

P F1 (pg 1 of 2) **Subatomic Particles: Protons, Neutrons, Electrons**

Name _____ Per _____

- Use the Periodic Table to fill in the chart below with the atomic number, mass number, number of protons, neutrons, and electrons. You may find NS C1 helpful for assistance with this practice.
- Remember that all atoms are neutral and have an equal number of protons and electrons.
- Remember that ions have lost or gained electrons (not protons).
- Unless otherwise told, assume the most common isotope, using the mass number from the periodic chart.

	symbol or name	atomic number	mass number	number protons	number neutrons	number electrons
1	carbon	6				
2	Ar		40			
3	iodine				74	
4	Na			11		
5	neon					
6	Ti ²⁺	22				
7	Cl ⁻			17		
8		19				18
9				15		18
10	¹⁰ Be					
11	¹⁴ O					
12	Indium-118			49		
13	⁶⁹ Zn					
14		30			34	
15	hydrogen-2					

	symbol	atomic number	mass number	number protons	number neutrons	number electrons
1	carbon	6	12	6	6	6
2	argon	18	40	18	22	18
3	iodine	53	127	53	74	53
4	Na	11	23	11	12	11
5	neon	10	20	10	10	10
6	Ti ²⁺	22	48	22	26	20
7	Cl ⁻	17	35	17	18	18
8	K ⁺	19	39	19	20	18
9	P ³⁻	15	31	15	16	18
<i>For problems 10-15, these symbols are indicating an isotope other than the average one listed on the periodic table.</i>						
10	¹⁰ Be	4	10	4	6	4
11	¹⁴ O	8	14	8	6	8
12	¹¹⁸ In	49	118	49	69	49
13	⁶⁹ Zn	30	69	30	39	30
14	⁶⁴ Zn	30	64	30	34	30
15	² H	1	2	1	1	1